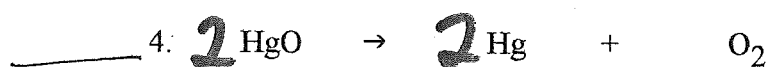
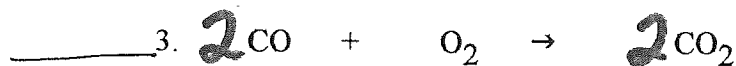
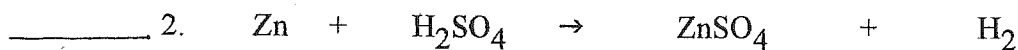
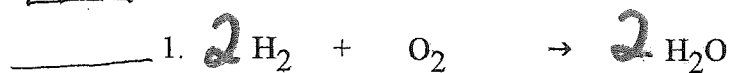


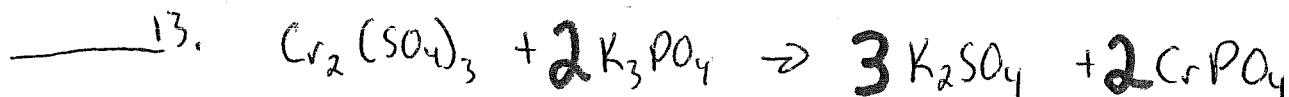
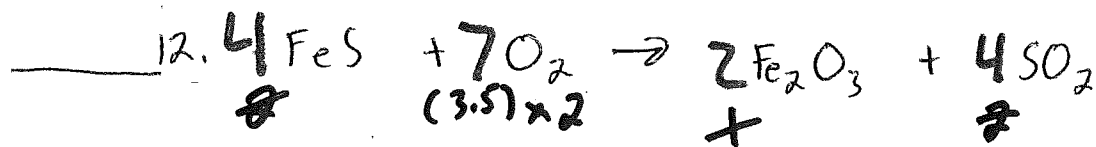
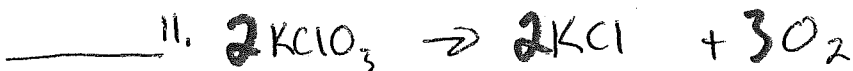
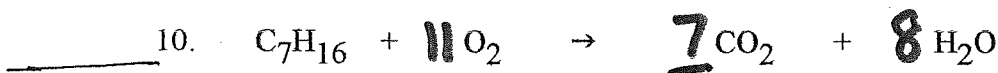
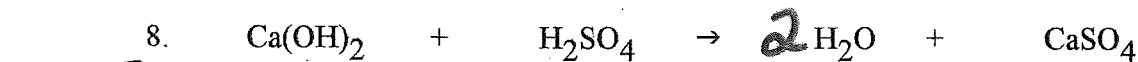
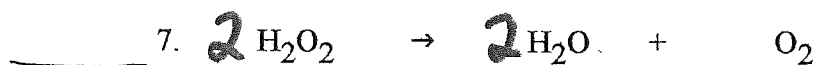
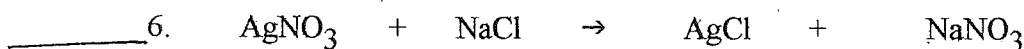
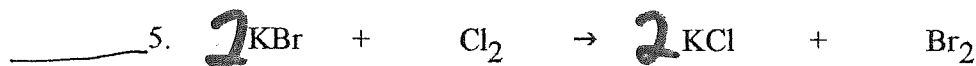
Types of Reactions

Classify the reactions below as combination, decomposition, single replacement, double replacement, or combustion. Then balance each equation.

Type



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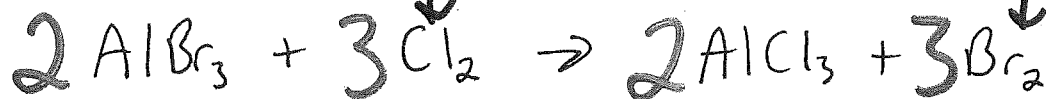


Write balanced formula equations for each reaction below.

1. zinc + lead (II) nitrate yield zinc nitrate + lead



2. aluminum bromide + chlorine yield aluminum chloride + bromine



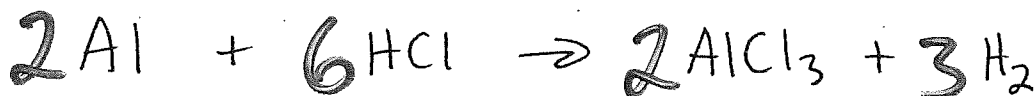
3. sodium phosphate + calcium chloride yield calcium phosphate + sodium chloride



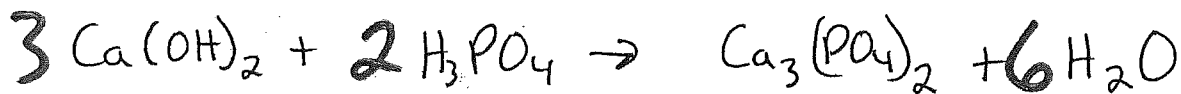
4. potassium chlorate when heated yields potassium chloride + oxygen gas



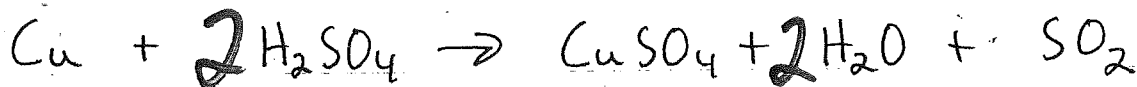
5. aluminum + hydrochloric acid yield aluminum chloride + hydrogen gas



6. calcium hydroxide + phosphoric acid yield calcium phosphate + water



7. copper + sulfuric acid yield copper (II) sulfate + water + sulfur dioxide



8. hydrogen + nitrogen monoxide yield water + nitrogen

